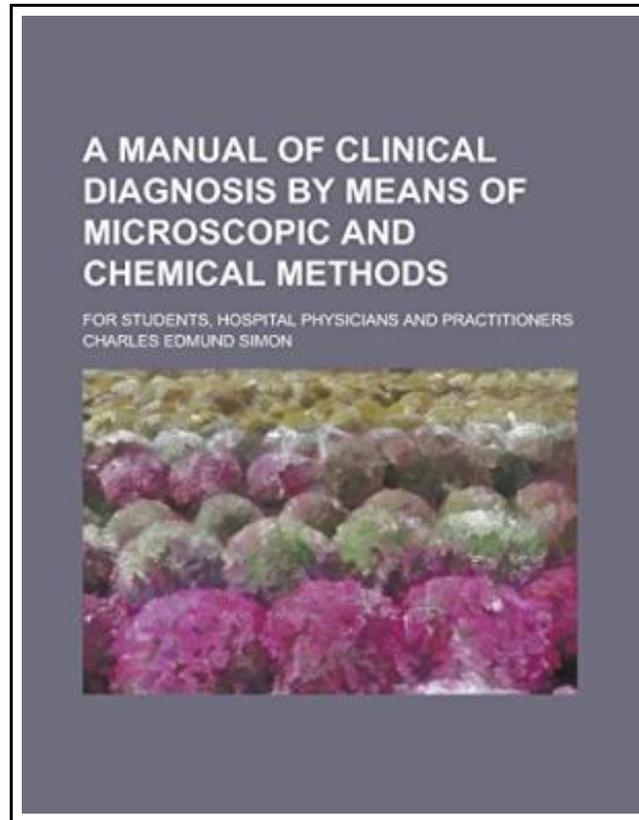


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Rarebooksclub.com, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1896 edition. Excerpt: .found. This figure is then corrected for the total amount of urine. Hence the equations, I., $x = 15 - 2$; II., 1000: $x :: A : Ch$, or in which Ch represents the quantity of chlorides contained in the total amount of urine, A the amount of urine actually passed, n the number of c.c. of the KSCN solution used in the precipitation of the excess of chlorides in 80 c.c. of the filtrate. $i 6.5 T = 7.05$. So in the above case $Ch = 600 1000$ The method described may be employed in the presence of albumin, albumoses, peptones, and sugar; the urine, however, must be fresh, so as to insure the absence of nitrous acid. Direct Method. If absolute accuracy is not required, the following method may be employed: Ten c.c. of urine are diluted with distilled water to 100 c.c. and treated with a few drops of a solution of potassium chromate. This mixture is titrated with a one-tenth normal solution of silver nitrate until the end-reaction--i. e., the occurrence of a faint orange tinge, which no longer disappears on stirring--is reached. The number of c.c. used multiplied by 0.01 will indicate the amount of chlorides present in 10 c.c. of urine. As uric acid, the xanthine bases, hypo-sulphites, sulpho-cyanides, and pigments are also precipitated by the silver nitrate, the endreaction is delayed; moreover, unless the urine be very pale, its recognition may be very difficult, and the error thus caused quite considerable....



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